

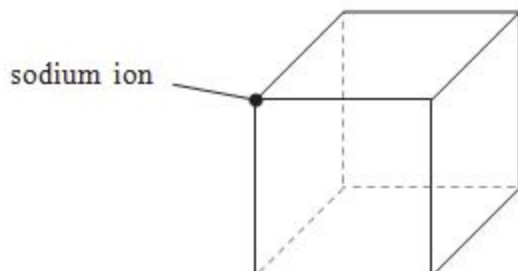
WJEC Chemistry A-level

C1.5: Solid Structures

Practice Questions

England Specification

1. (a) A section of the crystal structure of sodium chloride is shown below.



(i) Indicate, with a cross, the position of any chloride ion on this diagram

[1]

(ii) State the crystal co-ordination number of a **chloride** ion in the structure of sodium chloride.

[1]

(b) 'Rock salt', used on roads in winter, consists mainly of crystalline sodium chloride that is contaminated by a small quantity of insoluble mudstone.

Gwen added powdered rock salt to water and filtered out the insoluble material. She then evaporated the filtrate to dryness to produce pure white crystals of sodium chloride.

State **two** steps that she should have carried out to ensure that she obtained the **maximum** amount of sodium chloride from her rock salt crystals.

[2]

(c) The minerals 'rock salt', NaCl , and kainite, $\text{KCl} \cdot \text{MgSO}_4 \cdot 3\text{H}_2\text{O}$, both contain chloride ions.

(i) Give a chemical test that produces the same result for both of these compounds. You should state the reagent(s) used and the result of the test.

[2]

(ii) Give a chemical test, other than a flame test, that will show that these two compounds are different. You should assume that they are present as aqueous solutions.

Give the reagent(s) used and the result of the test for each compound.

[2]

(d) A common reaction of the halogens is the formation of the anion, X^- .

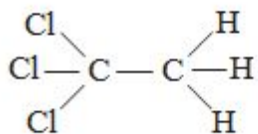
(i) State, in terms of electronic structure, why this occurs.

[1]

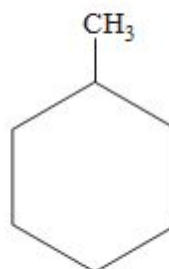
(ii) Give a reason why the tendency to form the X^- ion decreases down the halogen group.

[1]

(e) One compound previously used in correction fluid was 1,1,1-trichloroethane, but this has been replaced by compounds such as methylcyclohexane, which has a much less adverse effect on the environment.



1,1,1-trichloroethane

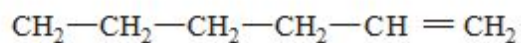


methylcyclohexane

(i) Explain, in terms of bond strengths, why 1,1,1-trichloroethane has an effect on the ozone layer but methylcyclohexane does not.

[2]

(ii) Hept-1-ene is an isomer of methylcyclohexane. CH_3-



Describe a chemical test that gives a positive result for hept-1-ene but not for methylcyclohexane. [2]

Reagent(s)

Observation

(Total 14)

2. (a) Both sodium chloride and caesium chloride have giant ionic structures.

(i) Draw a labelled diagram to show the arrangement of ions in a crystal of caesium chloride.

[2]

(ii) Give a reason why sodium chloride has a different structure from caesium chloride.

[1]

(b) Both diamond and graphite have giant covalent structures

(i) Describe the structure and bonding in graphite.

[3] QWC [1]

(ii) Explain why graphite can conduct electricity whilst diamond cannot.

[2]

(iii) Iodine, I_2 , also contains covalent bonds. Explain why solid iodine can be converted into a vapour at a much lower temperature than diamond. [3]

Total [12]

3. (a) Explain the fact that the melting temperature of sodium is much lower than the melting temperature of magnesium.

You should include reference to the type(s) of bonding involved and how this bonding affects melting temperatures. You may include a diagram if you consider it helpful. [3]

(b) In an experiment, 1-chlorobutane was heated with aqueous sodium hydroxide and the resulting solution was acidified. Aqueous silver nitrate was then added and a white precipitate was observed.

The experiment was repeated using 1-bromobutane and in this case a cream precipitate was observed. Explain these observations.

You should include:

- the type of reaction that occurs between the halogenoalkane and sodium hydroxide
- an equation for this reaction
- the identity of the coloured precipitates.
- an equation to show the formation of these precipitates.

[4] QWC [1]

(c) Describe how the structures of sodium chloride and caesium chloride are similar and how they are different. Give a reason for any difference. You may include a diagram if you consider it helpful.

[3]

(d) When hydrogen bromide, HBr, is added to propene, C₃H₆, two different products are possible. In practice, however, more of one of the products is formed. Explain why more of one product is formed.

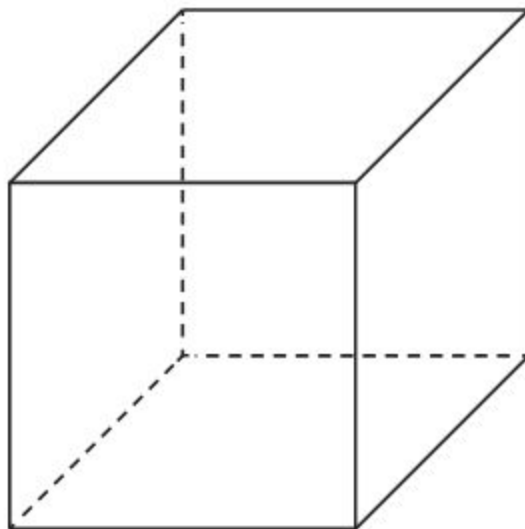
You should:

- state the type of reaction involved
- identify the two possible products
- state which of the two products predominates.
- give the reason why more of this product is formed.

[4] QWC [1]

(Total 16)

4. Complete and label the diagram to show the positions of the ions present in caesium chloride, CsCl. [1]



(Total 1)

5. When the temperature is increased, both solid iodine and diamond change directly into their gaseous state - they sublime.

(a) In each case, name the force or bond that is being overcome when the solid changes into a gas.

[2]

Iodine.....

Diamond.....

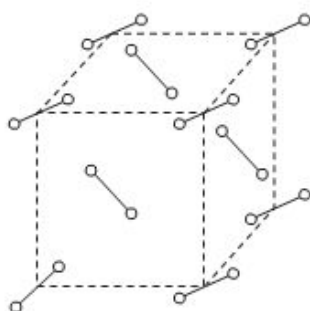
(b) State, with a reason, which solid would have the higher sublimation temperature.

[1]

(Total 3)

6.

(a) The structures of solid iodine and diamond are shown below.



Iodine



Diamond

Use these diagrams to help you explain why

- iodine vapourises easily but diamond does not vapourise until about 3550 °C
- neither iodine nor diamond conduct electricity

[4]
QWC [1]

